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Chapter 19 – Acids, Bases, and Salts.  
Jennie L. Borders. Section 19.1 – Acid-  
Base Theories. Acids have a sour taste,  
change the color of an indicator, can be  
strong or weak electrolytes in aqueous

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Answers and react with metals. Bases.

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Chapter 19“Acids, Bases, and Salts ...

Thus, a conjugate acid-base pair is related by the . loss or gain. of a single hydrogen ion. Chemical Indicators? They are weak acids or bases that have a different color

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Answers from their original acid and base. Acids and bases come in pairs.

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Chapter 19: Acids, Bases, and Salts.

STUDY. PLAY. Tastes sour. Acid.

Changes the color of an acid-base indicator (acid, base, or both) Both. Can

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~~Answers~~  
be strong or weak electrolytes in aqueous solution. Acid. In drinks, citrus, pop, etc. Acid. This acid is associated with protein. amino acid. Used in fragrances and flavors.

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44 terms. Taneya\_Heath. Chapter

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~~Answers~~ 19-Acids, Bases, and Salts. STUDY.

PLAY. Characteristics of Acids. they taste sour, will change the color of an acid-base indicator, and can be strong or weak electrolytes in an aqueous solution.  
Characteristics of Bases.

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When these acids and bases are mixed in the right proportions, the neutralization reaction thus results in the formation of salt and water. Some naturally occurring salts found in nature include NaCl and KCl etc in seawater and natural rock deposits.

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Acids: taste sour, react with certain metals to produce hydrogen gas, react with carbonates/bicarbonates to produce carbon dioxide gas. Corrode metals, pH is less

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Answers  
than 7, turns blue litmus paper red. Bases: taste bitter, have a slippery feeling. React with acids to form salts and water, pH is greater than 7.

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Blue litmus paper turns red in contact with

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an acid. Acids React with Carbonates and Bicarbonates.  $\text{HCl} + \text{NaHCO}_3 \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{CO}_2$ . Hydrochloric acid + sodium bicarbonate. salt + water + carbon dioxide. An old-time home remedy for relieving an upset stomach. OPERATIONAL BASES.

~~Chapter 19 Acids, Bases, and Salts~~

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~~MOLEBUS (ALLCHEM)~~

Chapter 19 Acids, Bases, and Salts. Acids and Bases: Basics. These are really just a specific type of chemical compound. No mysteries here. 1) They **MUST** be ionic compounds, and. **MUST** dissolve in water, that's the only way to be chemically active. We call this. disassociation.



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## ~~Chapter 19 Acids, Bases, and Salts~~

Acid + base  $\rightarrow$  salt + water + heat.  $\text{H}_2\text{SO}_4 + \text{MgO} \rightarrow \text{MgSO}_4 + \text{H}_2\text{O}$ .  $2\text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + 2\text{H}_2\text{O}$ . 2. Reaction of non-metal oxides with bases. Non-metal oxides are acidic in nature. Base + Non-metal oxide  $\rightarrow$  salt + water + heat.

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$2\text{NaOH} + \text{CO}_2 \rightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O}$ . To know more about Properties of Acids and Bases, visit here.

~~Class 10 Chapter 2 Science Notes Acids, Bases, and Salts~~

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fAcids Neutralize Bases  $\text{HCl} + \text{NaOH}$   
 $\text{NaCl} + \text{H}_2\text{O}$  -Neutralization reactions

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~~Answers~~ ALWAYS produce a salt (which is an ionic compound) and water. -Of course, it takes the right proportion of acid and base to produce a neutral salt fSulfuric Acid =  $H_2SO_4$

~~Chapter 19 Acids \_ Bases \_ and Salts | Acid  
| Ph | Free 30 ...~~

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Answers  
weak acid and one of its salts. or. a . weak  
base and one of its salts. Buffers. A buffer  
system is better able to resist changes in  
pH than pure water. Since it is a pair of  
chemicals: one chemical neutralizes any .  
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Potassium Chloride (KCl): It is formed

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after the reaction between potassium hydroxide (a strong base) and hydrochloric acid (a strong acid). (ii)

**Acidic Salts:** Salts which are formed after the reaction between a strong acid and weak base are called Acidic salts. The pH value of acidic salt is lower than 7.

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