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A solution is a homogeneous mixture of two or more components in which the particle size is smaller than 1 nm. Common examples of solutions are the sugar in water and salt in water solutions, soda water, etc. In a solution, all the components appear as a single phase. There is particle homogeneity i.e. particles are evenly distributed.

Solution - Definition, Properties, Types, Videos & Examples

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A solution is a homogeneous mixture of two or more components. The dissolving agent is the solvent. The substance that is dissolved is the solute. The components of a solution are atoms, ions, or molecules, making them 10⁻⁹ m or smaller in diameter.

Solutions, Suspensions, Colloids, and Dispersions

2 millimoles (mmol) = 2 × 10⁻³ mol (0.002 mol) centi: c: 10⁻²; 7 centimeters (cm) = 7 × 10⁻² m (0.07 m) deci: d: 10⁻¹; 1 deciliter (dL) = 1 × 10⁻¹ L (0.1 L) kilo: k: 10³; 1 kilometer (km) = 1 × 10³ m (1000 m) mega: M: 10⁶; 3 megahertz (MHz) = 3 × 10⁶ Hz (3,000,000 Hz) giga: G: 10⁹; 8 gigayears (Gyr) = 8 × 10⁹ yr (8,000,000,000 Gyr) tera: T: 10¹²

Measurements | Chemistry

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Solution, in chemistry, a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility. The term solution is commonly applied to the liquid state of matter, but solutions of gases and solids are possible.

solution | Definition & Examples | Britannica

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Solved: A rectangular block has dimensions 2.9 cm × 3.5 cm ...

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