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the subshell before pairing electrons.
2. packets/quanta of electromagnetic energy
3. the SI unit of frequency
4. An atomic orbital can hold no more

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Answers
than two electrons. 5. the amount of energy required to move an electron from its present energy level to the next higher one 6. the modern description of the location and energy of electrons in an atom 7.

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through a tube filled with neon gas....

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1) Electrons circulate around the nucleus 2) Electrons can jump from one level to the next but can NOT exist between levels (ladder) 3) Energy

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Answers are NOT evenly spaced
(escalator) 4) Energy lost or gained is
NOT always the same 5) The higher
the energy level, the easier for the
electrons to leave (magnet)

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Answers
the formula $2n^2$. the maximum number of electrons that can occupy an energy level. opposite spin. in order to occupy the same orbital 2 electron must have. filled energy sub levels. stable electron configurations are likely to contain. all the orbitals contain one electron, with spins parallel.

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Review Flashcards | Quizlet~~

states that single electrons with the same spin must occupy each equal-energy orbital before additional electrons with opposite spins can occupy the same orbitals. Pauli

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~~Answers~~ exclusion principle. states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite spins.
photoelectric effect.

~~Chemistry Vocabulary Chapter 5
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