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Q = m.c. T. Q glass = 1000.0.2. (90-20) = 14000 cal. Q water = 1000.1. (90-20) = 70000 cal. Q calorimeter = 70000 + 14000 = 84000 cal. 1 mol S is 32 g. Molar combustion enthalpy of S is 84000 cal or 84 kcal. Since it is combustion enthalpy; H Combustion S = -84 kcal/mol. Thermochemistry Exams and Problem Solutions < Prev.

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Thermochemistry Thermochemistry and Energy and Temperature Thermochemistry is study of changes in energy (heat) associated with physical or chemical changes. Force = push F = m a (mass x acceleration) force units: N (newton) = kg m s⁻² Work = force x distance W = F d energy units: J (joule) = kg m² s⁻²

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H = M x C x T. Where M is the mass or volume of the solution being used to test the energy change, C is the Specific Heat Capacity of the solution, (for example the specific heat capacity of water is 4.1855 J g⁻¹ K⁻¹), and T is the change in temperature.

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Thermochemistry is a branch of Thermodynamics, a scientific field that comprises the relationships between heat, work, and other forms of energy resulting from different chemical and physical processes. Thermochemistry itself is defined as energy changes when a chemical reaction takes place. Energy is the capacity to supply heat or do work.

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The answer is given per mole of reaction. fHo (298 K) for SF₄ (g), H₂O (l), SO₂ (g) and HF (g) are -763, -286, -297 and -273 kJ mol⁻¹, respectively. What is the value of rHo (298 K) for the hydrolysis shown above? For this process, cHo (298 K) = -3953 kJ mol⁻¹.

~~Chapter 2: Thermochemistry~~

42. Given that heat neutralization of strong acid and strong base is -13.7 kcal. The heat produced when one mole of HCl is mixed with 0.5 mole of NaOH will be ____